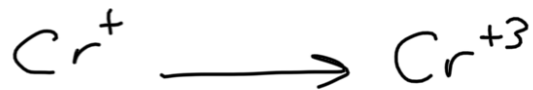


Saved Photo

1. Identify the species being oxidized and reduced in each of the following reactions (give detailed reasoning):

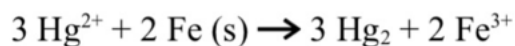


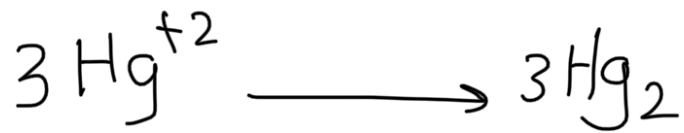
Cr oxidized because oxidation number increases $+1 \rightarrow +3$



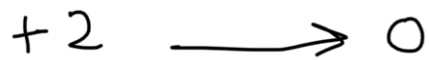
Sn reduced because oxidation number decreases

2. Identify the species being oxidized and reduced in each of the following reactions (give detailed reasoning):





Hg reduced because oxidation number decreases



Fe oxidized because oxidation number increases

